

Module 2: Biochemistry

Topic 2 Content: Acids and Bases Notes



An acid is a solution that has an excess of H^+ ions. The more H^+ ions a substance has, the more acidic the solution becomes. Properties of an acid include a low pH, a sour taste, electrical conductivity, corrosion of metals, and turning blue litmus paper red. There are many uses of acids. Citric acid is found in fruits, such as lemons, limes, and oranges. Ascorbic acid is a key acid which your body needs to function. It is found in vitamin C. Sulfuric acid is used in the production of fertilizers, steel, paints, and plastics.



A base is a solution that has an excess of OH^- ions. Bases are substances that can accept hydrogen ions. Properties of a base include; a high pH, feeling slippery, tasting bitter, corroding metals, conducting electricity, not reacting with metals, and turning red litmus paper blue. There are also many uses to bases. The OH^- ion in bases interacts with substances, such as dirt and grease. When bases are added to cleaning products, they are more efficient. One of the most important fluids in your body is basic: your blood.